LET,* Department of Chemistry, University of Minnesota, Minneapolis, Minnesota 55455.

On p 7083 the ligand dppp is named incorrectly and should be 1,3-bis(diphenylphosphino)propane, not 1,2-bis(diphenylphosphino)propane.

Sensitivity Enhancement in Natural Abundance Proton-Coupled ^{15}N NMR Spectra Using the Selective Population Transfer Method [*J. Am. Chem. Soc.*, **101**, 774 (1979)]. By HANS J. JAKOBSEN* and WALLACE S. BREY*, Department of Chemistry, University of Florida, Gainesville, Florida 32611.

In the caption for Figure 1, values for two of the coupling constants were interchanged. The correct assignments, with values determined by computer fitting of the ^{15}N spectrum of pyrrole to yield an rms error of 0.010 Hz, are: $^2J_{^{15}N\text{-}H(2)} = -4.53$ Hz and $^3J_{^{15}N\text{-}H(3)} = -5.35$ Hz. In the last paragraph of the text, the third sentence should conclude "until acquisition" rather than "during acquisition".

Methyl-Substituted Allyl Cations. A Comparison of Experimental Stability, Rotational Barrier, and Solvolysis Data with ab Initio Calculations [*J. Am. Chem. Soc.*, **101**, 6032 (1979)]. By HERBERT MAYR,* WOLFGANG FÖRNER, and PAUL VON RAGUÉ SCHLEYER, Institut für Organische Chemie der Universität Erlangen-Nürnberg, 8520 Erlangen, West Germany.

Page 6035, eq 10: C_2H_8 should be C_3H_8 and C_3H_6 should be C_2H_6.

Free-Energy Relationships for Electron-Transfer Processes [*J. Am. Chem. Soc.*, **101**, 6140 (1979)]. By FRANCO SCANDOLA,* Centro di Studio sulla Fotochimica e Reattività degli, Stati Eccitati dei Composti di Coordinazione del CNR, University of Ferrara, Ferrara, Italy, and VINCENZO BALZANI,* Istituto Chimico "G. Ciamician" dell'Università and Laboratorio di Fotochimica e Radiazioni, d'Alta Energia del CNR, Bologna, Italy.

Two footnotes (refs 45 and 46) are lacking. The footnotes are as follows:

(45) In this view, the α and β coefficients of eq 4 can be given as a function of ΔG and $\Delta G^{\ddagger}(0)$ as follows:

$$\alpha = 1/[1 + \exp(-\ln 2\ \Delta G/\Delta G^{\ddagger}(0))]$$

$$\beta = \Delta G/[1 + \exp(\ln 2\ \Delta G/\Delta G^{\ddagger}(0))] + \Delta G^{\ddagger}(0) \ln [1 + \exp(-\ln 2\ \Delta G/\Delta G^{\ddagger}(0))]/\ln 2$$

(46) For an example of such behavior, see: I. B. Afanas'ev, S. V. Prigoda, T. Y. Mal'tseva, and G. I. Samokhvalov, *Int. J. Chem. Kinet.*, **6**, 643 (1974).

Conformational Characteristics of Poly(methyl vinyl ketone)s and of Simple Model Ketones [*J. Am. Chem. Soc.*, **101**, 6481 (1979)]. By ULRICH W. SUTER, the Eidgenössische Technische Hochschule, Technisch-Chemisches Laboratorium, ETH-Zentrum, CH-8092 Zürich, Switzerland.

On page 6482, on the next to the last line in the formula, $f_{6\text{-}12}(r_r)$ should be replaced by $f_{6\text{-}12}(r^*)$, and the same substitution should be made on page 6483 (left-hand column), line 4.

Structure and Reactivity in Intramolecular Catalysis. Catalysis of Sulfonamide Hydrolysis by the Neighboring Carboxyl Group [*J. Am. Chem. Soc.*, **101**, 6981 (1979)]. By TEUN GRAAFLAND, ANNO WAGENAAR, ANTHONY J. KIRBY, and JAN B. F. N. ENGBERTS,* Department of Organic Chemistry, University of Groningen, Nijenborgh, 9747 AG Groningen, The Netherlands, and the University Chemical Laboratory, Cambridge, CB2 1EW, England.

In Table III the thermodynamic activation parameters for hydrolysis of **11** in 50% EtOH–H_2O (v/v) should be: $\Delta G^{\ddagger} = 26.39$ kcal mol^{-1} and $\Delta S^{\ddagger} = -18$ eu. $\Delta H^{\ddagger}$ is given correctly. The conclusions are not affected.

A New Empirical Method to Calculate Average Molecular Polarizabilities [*J. Am. Chem. Soc.*, **101**, 7206 (1979)]. By KENNETH J. MILLER* and JOHN A. SAVCHIK, Department of Chemistry, Rensselaer Polytechnic Institute, Troy, New York 12181.

Page 7208, second column: Equation 21 should read:

$$\sqrt{\overline{r_A^2}} = \sqrt{\frac{3}{2}}\sqrt[4]{\overline{a_0\alpha_A}} \qquad (21)$$

Intramolecular Electrostatic and General Acid Catalysis in the Hydrolysis of *O,S*-Thioacetals [*J. Am. Chem. Soc.*, **102**, 292 (1980)]. By THOMAS H. FIFE* and THEODORE J. PRZYSTAS, Department of Biochemistry, University of Southern California, Los Angeles, California 90033.

Page 296, column 2, line 28: "C–O bond breaking" should read "C–S bond breaking".

Photolysis of Alkyl Azides. Evidence for a Nonnitrene Mechanism [*J. Am. Chem. Soc.*, **102**, 735 (1980)]. By EVAN P. KYBA* and RUDOLPH A. ABRAMOVITCH,* Departments of Chemistry, The University of Texas at Austin, Austin, Texas 78712, The University of Alabama, University, Alabama 35486, and Clemson University, Clemson, South Carolina 29631.

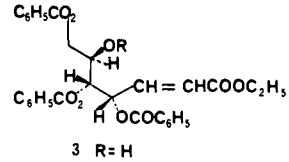
Page 740: In the Acknowledgments, reference 33 should appear after the initials R. A. A. inside the parentheses and *not* outside.

The Chemistry of Bis(fulvalene)divanadium [*J. Am. Chem. Soc.*, **102**, 1009 (1980)]. By JAMES C. SMART* and BARRY L. PINSKY, Department of Chemistry, University of California, Berkeley, California 94720.

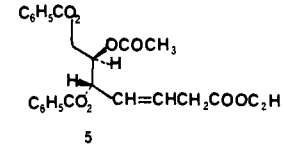
On page 1010, column 1, line 6 should read, "Bis(fulvalene)divanadium(II,III) Hexafluorophosphate, $[(\eta^5\text{:}\eta^5\text{-}C_{10}H_8)_2V_2^+](PF_6^-)$."

Stereospecific Total Synthesis of a "Slow Reacting Substance" of Anaphylaxis, Leukotriene C-1 [*J. Am. Chem. Soc.*, **102**, 1436 (1980)]. By E. J. COREY*, DAVID A. CLARK, GIICHI GOTO, ANTHONY MARFAT, and CHARLES MIOSKOWSKI, Department of Chemistry, Harvard University, Cambridge, Massachusetts 02138, and BENGT SAMUELSSON and SVEN HAMMARSTROM, Karolinska Institutet, S-10401 Stockholm, Sweden.

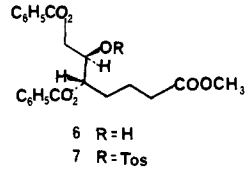
On p 1437 structures **3–9** were incorrectly drawn. The correct structures are shown below.



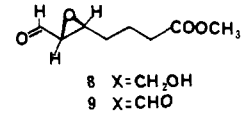
3 R = H
4 R = OAc



5



6 R = H
7 R = Tos



8 X = CH_2OH
9 X = CHO